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the equilibrium constant  $K_p$  for the following reaction is  $1.07 \times 10^{-7}$  at  $521 \text{ K}$   $\text{SO}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons \text{SO}_2\text{Cl}_2(\text{g})$  calculate the equilibrium partial pressures of all species when  $5 \text{ g}$  of  $\text{SO}_2$  is introduced into an evacuated flask at a pressure of  $1.71 \text{ atm}$  at  $521 \text{ K}$  there are 3 steps to solve this one

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